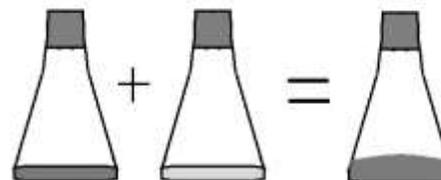


Explain the conservation of mass.

In a chemical reaction, the mass of the final product is the same as the mass of the reacting substances because atoms are rearranged but not destroyed.



Example:

When coal (carbon) is burned in the presence of air (oxygen), another gas (carbon dioxide) is formed.



Mark the statements **TRUE** or **FALSE**.

When the coal was burned, the carbon was destroyed.

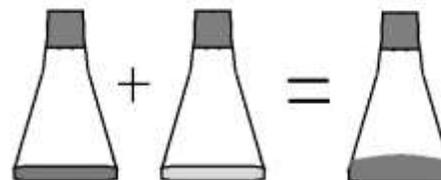
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The conservation of mass only applies to solids.

The number of atoms always remains the same.

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The number of atoms always remains the same.

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Weigh a small disposable cup with ¼ cup of water. Also weigh an Alka-Seltzer tablet. Add the weights. Now, drop the tablets into the water and watch what happens. When the chemical reaction is complete, weigh the cup. It is lighter than the sum of the two original substances. Now Repeat the experiment, but with the substances in a closed Ziploc bag that captures the gas. Explain that even though substances changed, no atoms were lost.

Chemistry 11 Information Pieces

True C-11
True C-11
False C-11
False C-11
C C-11
O C-11
CO₂ C-11

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